

**2. Correct. The answer is true.** The combustion of 'sour' hydrocarbons (e.g. methane that contains sulfur  $4\text{CH}_4 + \text{S}$ ) releases sulfur dioxide ( $4\text{CH}_4 + \text{S} + 10\text{O}_2 \Rightarrow 5\text{SO}_2 + 4\text{CO}_2 + 2\text{H}_2\text{O} + \text{energy}$ ). This sulfur dioxide in contact with water and oxygen in the atmosphere produces sulfurous acid ( $\text{SO}_2 + \text{H}_2\text{O} = \text{H}_2\text{SO}_3$ ) and sulfuric acid ( $2\text{SO}_2 + 2\text{H}_2\text{O} + \text{O}_2 = 2\text{H}_2\text{SO}_4$ ). These acids fall with the rain and can cause damage to fish, vegetation, forests, and buildings.