

2. Incorrect. The answer is true not false. The combustion of 'sour' hydrocarbons (e.g. methane that contains sulfur = $4\text{CH}_4 + \text{S}$) releases sulfur dioxide ($4\text{CH}_4 + \text{S} + 10\text{O}_2 \Rightarrow 5\text{SO}_2 + 4\text{CO}_2 + 2\text{H}_2\text{O} + \text{energy}$). This sulfur dioxide in contact with water and oxygen in the atmosphere produces sulfurous acid ($\text{SO}_2 + \text{H}_2\text{O} \Rightarrow \text{H}_2\text{SO}_3$) and sulfuric acid ($2\text{SO}_2 + 2\text{H}_2\text{O} + \text{O}_2 \Rightarrow 2\text{H}_2\text{SO}_4$). These acids fall with the rain and can cause damage to fish, vegetation, forests, and buildings.